Chemistry Reference Sheet

***Constants***

**Volume of 1 mole at STP** = 22.4 L

**Gas Constant (R)** = 0.0821 L\*atm/mole\*k or 8.314 L\*kPa/mole\*K

**STP** = 0 C at 1 atm or 273 k at 101.3 kPa

Pressures = 1 atm = 14.7 psi = 101.3 kPa = 760 torr = 760 mm Hg

**1 mole** = 6.02 x 1023 molecules

***Equations***

**Density** $D=\frac{m}{v}$

K = C +273

F⁰ = (1.8 X C⁰) + 32

C⁰ = (F - 32⁰) / 1.8

**Combined Gas Law** P1V1/T1 =P2V2/T2

**Ideal Gas Law** PV = nRT

**n** = moles