Name Period Date

#### Stoichiometry – Ch. 12

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| 1. How many grams of hydrogen gas would be needed to form 8.0 grams of water?

\_\_\_\_ H2 + ­­­\_\_\_\_ O2 → \_\_\_\_ H2O |
| 1. How many molecules of oxygen are produced when 29.2 g of water is decomposed?

\_\_\_\_ H2O → \_\_\_\_ H2 +\_\_\_\_O2 |
| 1. Assuming STP, how many mL of oxygen are needed to produce 20.4 mL SO3.

\_\_\_\_SO2 + \_\_\_\_O2 → \_\_\_\_SO3 |
| 1. How many Liters of oxygen are required to burn 3.86 L of CO2? How many Grams?

2CO2 + O2 → 2CO2 |